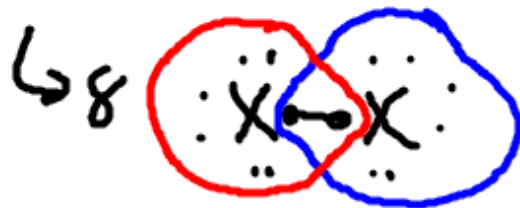


Lewis Structures

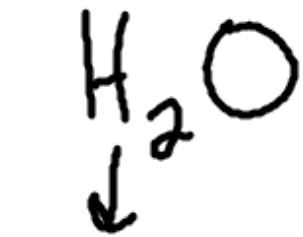
for molecules

for now will duet/octet rule



Rules/Steps

- 1) determine total number of val e^- present
- 2) attach all atoms with single bonds
- 3) add remaining e^- so they follow octet rule



$$(1 \times 2) + 6 = 8e^-$$

$$\frac{-4}{4e^-}$$

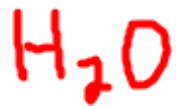


unshared pr

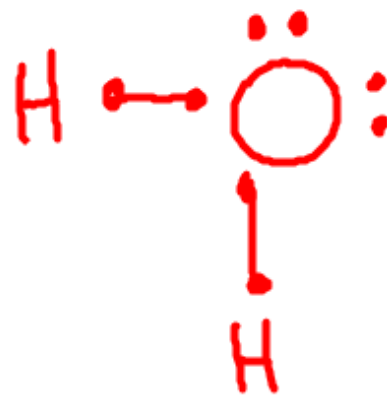
shared pr



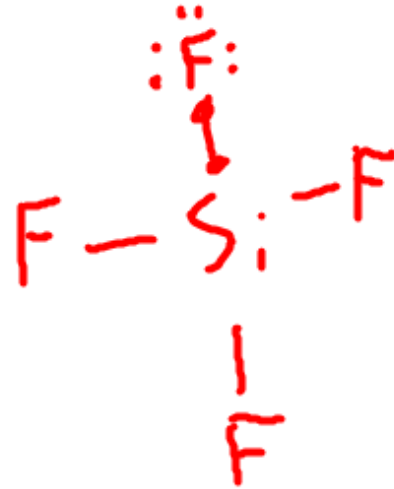
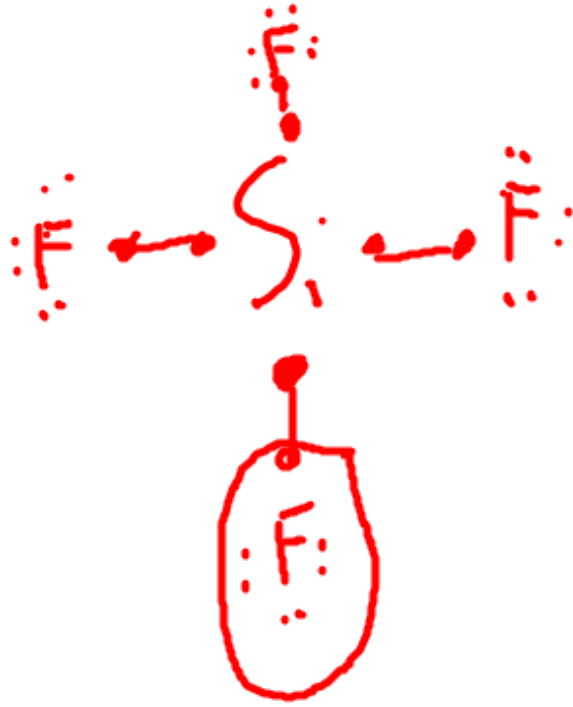
other way



- 1) central with val e⁻
- 2) Attach outsiders
- 3) count

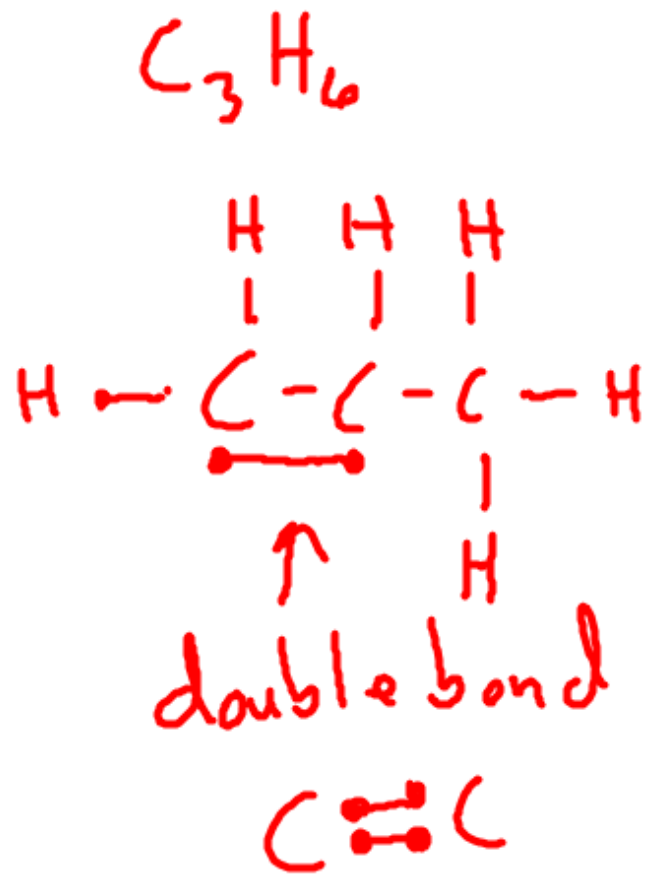
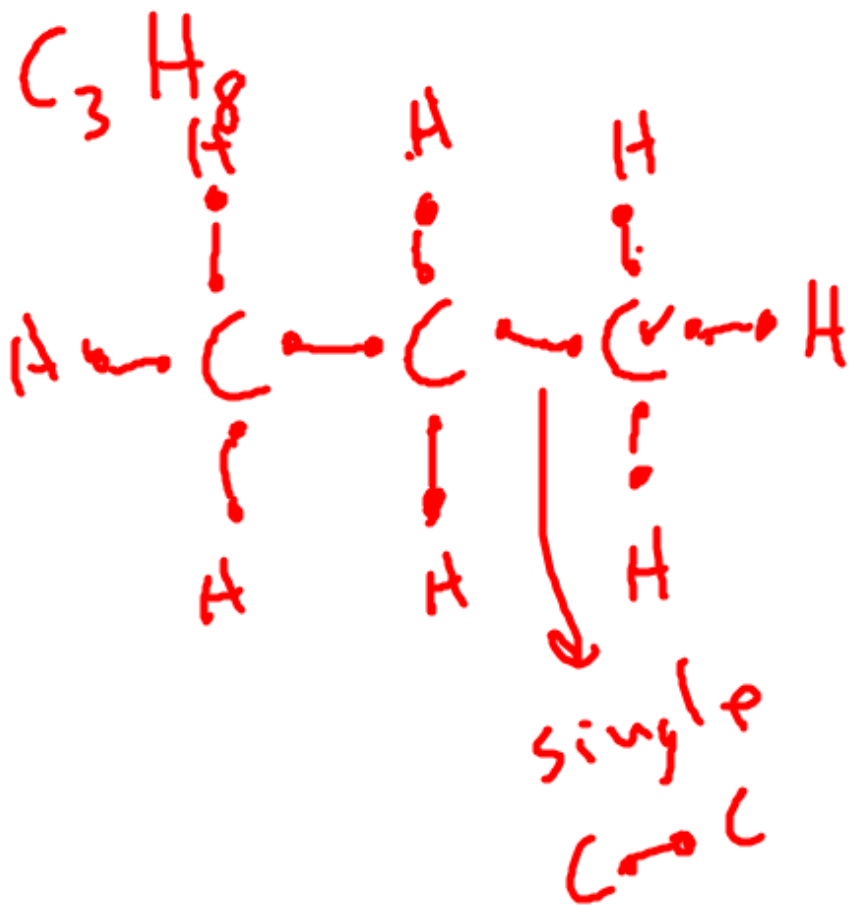


$$\begin{array}{cc}
 \text{Si} & \text{F}_4 \\
 \downarrow & \downarrow \\
 4 & + 7 \cdot 4 = \frac{32}{8} \\
 & \underline{\quad} \\
 & 24
 \end{array}$$

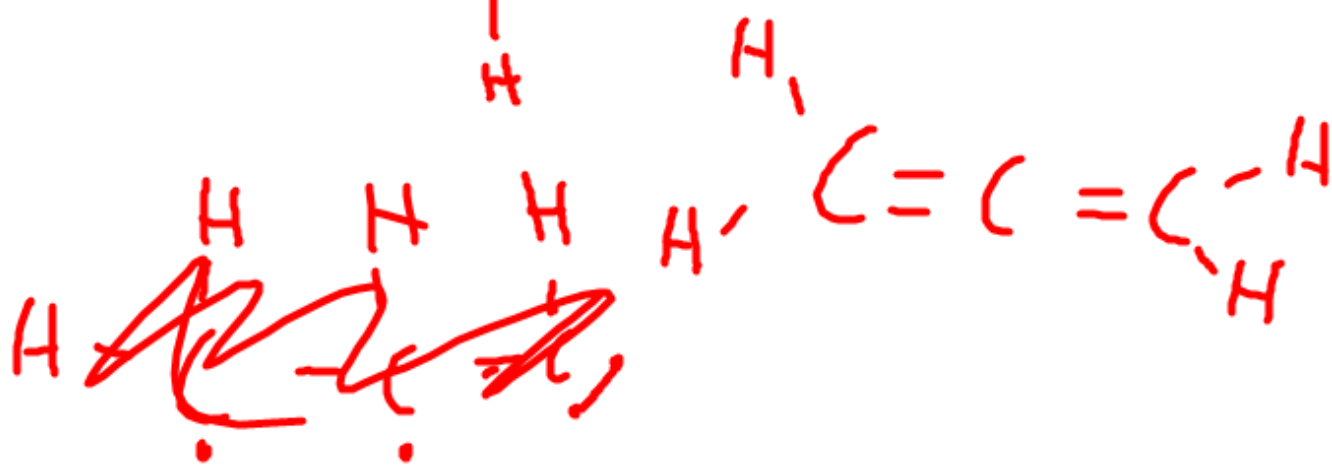
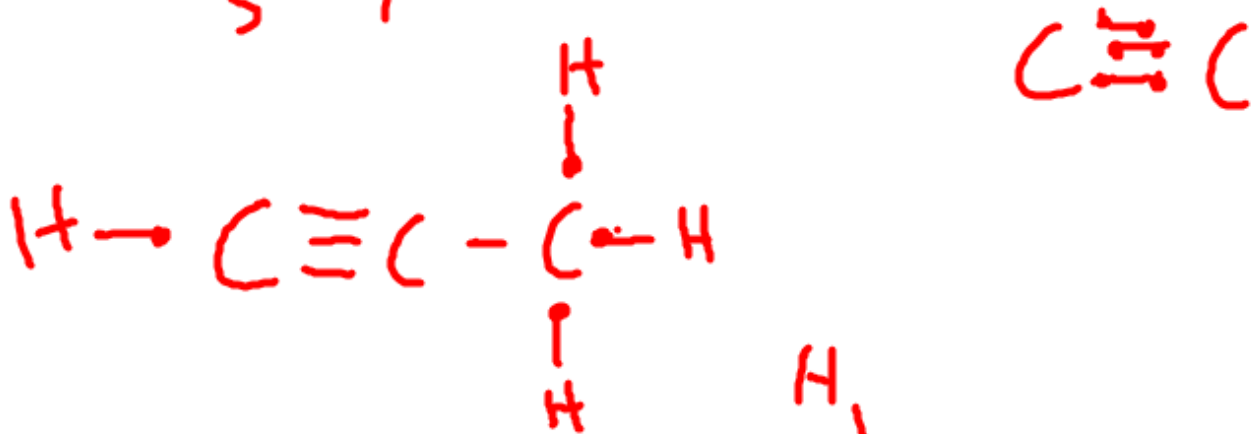


hydrocarbon - H & C

1) attach C together



C_3H_4 triple bond



ions

